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METHOD OF REMOVING CALCIUM FROM WATER COUNTAINING  
CALCIUMHYDROGEN CARBONATE IN HIGH CONCENTRATION

**Field of the Art:**

5 The present invention relates to a method for removing calcium from water containing a high concentration of calcium bicarbonate.

More specifically, this invention relates to a method for removing calcium from defluorinated waste water by making the  
10 waste water containing fluorine react with calcium carbonate.

In semiconductor device manufacturing plants, a large quantity of water such as ultrapure water or the like is used for washing in addition to the agents such as hydrofluoric acid, ammonium fluoride. Reuse of the water by recovery and is an  
15 important problem.

**Background of the Art:**

Fluorinated agents consumed in a semiconductor device manufacturing plant are wasted with a large quantity of wash  
20 water.

In general, this waste water is processed with hydrated lime.

Namely, fluoride is wasted as poorly-soluble calcium fluoride, and water is drained out in a 15ppm or less  
25 concentration of fluorine.

Recovery of fluoride has been examined from the viewpoints of effective resource recovery and waste reduction.

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In order to recover fluorine as calcium fluoride of reusable quality, a conventional process by neutralization with hydrated lime results in producing too fine particles of calcium fluoride to separate, handle, and dry them, and the process is 5 too difficult to be put into practice.

In order to obtain calcium fluoride of reusable quality as one of resources, it is necessary to use calcium carbonate as a source of calcium. Calcium carbonate reacts with fluorine and is converted into calcium fluoride without change in its 10 original shape and size. Namely, by using calcium carbonate having a  $50\text{ }\mu\text{m}$  average particle diameter, calcium fluoride of approximately a  $50\text{ }\mu\text{m}$  average particle diameter can be obtained, and this process operates with good efficiencies in dehydration and separation.

15 By removing fluorine by means of the above-mentioned neutralization method using calcium carbonate, the fluorine in the drain can be reduced to approximately 5ppm and cleaned treated water can be obtained.

Moreover, according to this neutralization method, 20 calcium is removed as  $\text{CaF}_2$ . Also, since excess calcium carbonate precipitates as it is, it has been considered that Ca is not contained in the drain water after treatment.

Nevertheless, through detail examinations of this treated water, it has been found out that there is a quite large amount 25 of calcium remained in the treated water. Namely, it has been found out that there is as much calcium as exceeding 75ppm in alkalinity ( $\text{CaCO}_3$ ) and 120ppm in hardness ( $\text{CaCO}_3$ ) specified as

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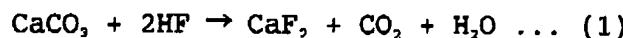
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the water quality standards for industrial use.

For this reason, as a result of zealous search into the cause, the applicants of the present invention have considered that calcium remains therein for the following reason. Namely,

5 carbon dioxide is by-produced by the reaction between calcium carbonate and hydrofluoric acid as shown in the following reaction formulae.



Carbon dioxide produced here is dissolved in water to 10 produce a carbonic acid.



It is considered that calcium bicarbonate is relatively soluble, therefore, calcium remains in the drain water after treatment.

15 The applicants of the present invention have confirmed that after de-fluorinating waste water of about 500 - 1000ppm fluorine concentrations with calcium carbonate, it contains about 200 - 500ppm calcium dissolved therein.

Thus, water of high calcium concentrations cannot be reused 20 even as cooling water as it is, because it causes troubles such as scaling.

It is necessary to reduce its alkalinity ( $\text{CaCO}_3$ ) and hardness ( $\text{CaCO}_3$ ) to about 75ppm and 120ppm specified by the water quality standards for industrial use.

25 On the other hand, as methods for reducing the alkalinity and hardness, various methods such classical methods as a degassing process, distillation process, an activated carbon

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absorption process, etc., and a ion-exchanger membrane process, dialysis membrane process, etc. have been provided.

However, it is limited from viewpoints of a life of equipment and its running costs to treat a large quantity of 5 water of calcium concentrations as high as 400 - 500ppm.

This purpose of the present invention is to provide a method for removing calcium from water containing a high concentration of calcium bicarbonate, permitting to reduce calcium bicarbonate equivalent to 200 - 500ppm as calcium to the level 10 in accordance with the water quality standards for industrial use by means of a simple chemical treatment and not by means of using much heat and power for heating and deaeration.

**Disclosure of the Invention :**

This invention provides a method for removing calcium from 15 water containing a high concentration of calcium bicarbonate, characterized in removing calcium by adding calcium hydroxide to waste water containing a high concentration of calcium formed as calcium bicarbonate, making the waste water react with it, and immobilizing calcium as calcium carbonate.

**20 Actions :**

The applicants of the present invention have tested and examined the method for reducing the dissolved calcium bicarbonate not by using such physical methods as heating and deairing, but by using, what is called, a chemical method.

25 The calcium dissolved as calcium bicarbonate can be removed by converting it into a poorly-soluble calcium salt.

As poorly-soluble salts, calcium fluoride (CaF<sub>2</sub>), calcium

carbonate ( $\text{CaCO}_3$ ), hydroxy-apatite ( $\text{Ca}_{10}(\text{PO}_4)_6(\text{OH})_2$ ), etc. can be mentioned.

Each solubility at a normal temperature is as shown in Table 1, and any of them is less than 10ppm as concentrations of Ca.

5 Table 1

Kinds of Salt	Solubility	Ca concentration (ppm)
Calcium fluoride	0.016g/l	8
Calcium carbonate	0.013g/l	5
Apatite	$\text{Ca}^{2+}$ 0.000123M/l	5

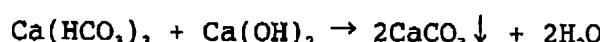
As a method for converting calcium bicarbonate dissolved in water into a poorly-soluble calcium salt, the following reactions can be mentioned, and each method has been examined.

10 (1) A method for removing calcium as calcium fluoride :



In this reaction,  $\text{CaF}_2$  is precipitated and Ca concentration is reduced. If hydrofluoric acid is added too much, the water under treatment may be increased in F concentration. In such 15 a case, it is necessary to remove fluorine by adding a proper quantity of calcium hydroxide.

(2) A method for removing calcium as calcium carbonate :



Calcium carbonate is precipitated by adding calcium 20 hydroxide. In this reaction, coexisting  $\text{HCO}_3^-$  is also removed as calcium carbonate

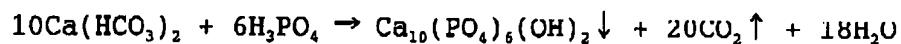
at the same time, therefore, it is very reasonable for purifying the water.

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(3) A method for removing calcium as apatite :



Hydroxy-apatite is precipitated by adding phosphoric acid. 5 however, this reaction is considered to be conditional on a hydrogen ion concentrations (pH) or the like, and it is difficult to select the conditions, and the examinations by the applicants of the present invention did not result in almost anything like precipitation.

From the above, as a method for fixing and removing calcium 10 dissolved as calcium bicarbonate, such a method is most reasonable, as calcium is fixed and removed as calcium carbonate having a very poor solubility by adding calcium hydroxide thereto.

This reaction does not require heating or cooling, but can 15 sufficiently be achieved at an ordinary temperature.

The reaction is performed while stirring, and either one of an palindromic system and a continuous system will do. A schematic diagram of a continuous treatment flow is shown in Fig.1.

20 Since the reaction is completed instantaneously, it is not necessary to take a long time for it but only 30 minutes are enough. In the case of the continuous system, two treating tanks having a residence time of about 30 minutes may be cascaded for preventing a short pass.

25 Also in the case shown in Fig.1, a 1st treating tank 2 and a 2nd treating tank 3 are arranged in series. A waste water tank 1 is coupled to the 1st treating tank 2 via a flow controller

5. The waste water in the waste water tank 1 flows out of the bottom of the waste water tank 1, and flows into 1st treating tank from the top of it. Calcium hydroxide is added into the 1st treating tank 2 from outside. The liquid overflowing from 5 the 1st treating tank 2 flows into the 2nd treating tank 3 from the top of it. A precipitation tank 4 is coupled downstream from the 2nd treating tank 3. The liquid overflowing from the 2nd treating tank 3 flows into the precipitation tank 4, and solid-liquid separation is performed in the precipitation tank 10 4. The liquid overflowing the precipitation tank 4 is supplied for reuse as supernatant liquid. Slurry containing solids precipitated in the precipitation tank 4 is forwarded to a dehydrating separator.

The treating tanks 2, 3 each are provided with respective 15 stirrers 6, 7, which stir and mix up the liquid, the waste water, to be treated, and calcium hydroxide.

The products of the reaction are separated into solids and liquid according to a conventional method. Prior to the solid-liquid separation, it is preferable to concentrate the 20 precipitates by using the precipitation tank, a thickener, or the like. Moreover, in this case, the equipment can do with a smaller capacity by accelerating the precipitation using a small quantity of a polymer coagulant. For solid-liquid separation, a conventional use dehydrating separator such as 25 a centrifugal separator machine, a vacuum filter, a filter press are used.

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**Brief description of the Drawing:**

Fig.1 is a schematic flow diagram of a continuous treatment in accordance with the present invention.

**(Description of the Marks)**

5 1: Waste water tank, 2: 1st treating tank, 3: 2nd treating tank, 4: precipitation tank, 5: Flow controller, 6, 7 : Stirrer. 8: Calcium hydroxide adding equipment, and 9: pH meter.

**Best Mode for Embodying the Invention :**

The method of the present invention is disclosed in details 10 by showing embodiments in the following.

**(Embodiment 1)**

Treated water of pH = 7 containing 4ppm F and 240ppm Ca has been obtained by making waste water containing 5000ppm HF pass through a calcium carbonate packed tower. When one liter 15 of this treated water was sampled and a calcium reduction effect was examined varying an adding rate of calcium hydroxide, the results were obtained as shown in Table 2.

**(Table 2)**

Adding rate of calcium hydroxide	pH of treated water	Concentration of Ca (ppm)	XRD of solids obtained
50% of equivalent wt.	7.9	110	coincides with peak of $\text{CaCO}_3$
75% of equivalent wt.	8.3	56	coincides with peak of $\text{CaCO}_3$
100% of equivalent wt.	10.0	18	coincides with peak of $\text{CaCO}_3$
125% of equivalent wt.	11.1	66	coincides with peak of $\text{CaCO}_3$

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150% of equivalent wt.	11.7	97	coincides with peak of $\text{CaCO}_3$ and a quantity of $\text{Ca(OH)}_2$ , coexists
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The best calcium reduction effect could be obtained when an adding rate of calcium hydroxide was equivalent to the calcium concentration.

5 The solid obtained was calcium carbonate ( $\text{CaCO}_3$ ).

(Embodiment 2)

Drain water, which was obtained by making waste water containing 5000ppm HF pass through a packed tower with a 1.5m calcium carbonate layer height and defluorinating the waste water, exhibited  $\text{pH} = 7.5$ ,  $\text{F} = 4\text{ppm}$ , and  $\text{Ca} = 270\text{ppm}$ . From such an analysis as one liter of this drain water was sampled; 0.53 (Ca equivalent) calcium hydroxide of 95% purity was added to the water; it was treated for 30 minutes; and solid-liquid separated and analyzed, the results were obtained as shown in

15 Table 3.

Table 3

pH of treated water	Concentration of Ca (ppm)	Amount of solids obtained (g)	XRD of solids obtained
9.9	21	1.20	coincides with peak of $\text{CaCO}_3$

(Embodiment 3)

10 liters of drain water exhibiting  $\text{pH} = 7.5$ ,  $\text{F} = 4\text{ppm}$ ,  
20 and  $\text{Ca} = 240\text{ppm}$  by making waste water containing 5000ppm HF pass

through a packed tower with calcium carbonate of 1.5 layer height.

While making this drain water pass through the 1st tank of one liter capacity and the 2nd tank of two liters capacity 5 arranged in series as shown in Fig.1 at a treating rate of two liters/hour, the drain water was continuously treated for 5 hours by adding 0.23g (Ca equivalent) calcium hydroxide of 95% purity into the 1st tank 2 at 30 minute intervals, and analyzed and measured by sampling the treated water overflowing from 10 the 2nd treating tank 3 at every prescribed hours, to obtain the results as shown in Table 4.

(Table 4)

Elapsed time	pH	Concentration of Ca (ppm)	XRD of solids obtained
2 hours	8.9	22	coincides with peak of $\text{CaCO}_3$
3 hours	10.1	18	coincides with peak of $\text{CaCO}_3$
4 hours	9.8	23	coincides with peak of $\text{CaCO}_3$

Ca was stably reduced to 30ppm or smaller also by a 15 continuous treatment system, and cleared the water quality standards.

(Embodiment 4)

Waste water containing 5000ppm HF was defluorination-treated by making the water pass through a packed tower with 20 calcium carbonate. The analysis of the drain water exhibited pH = 6.8, F = 3ppm, and Ca = 480ppm.

After one liter of this drain water was stirred for 30

minutes by adding 1.40g calcium hydroxide of 95% purity thereto, the drain water was separated into solids and liquid, then the analysis exhibited the results as shown in Table 5.

(Table 5)

PH	F (ppm)	Ca (ppm)	XRD of solids obtained
10.2	2	27	coincides with peak of $\text{CaCO}_3$

5

(Embodiment 5)

Four liters of waste water treated through the calcium carbonate packed tower exhibiting pH = 6.8, F = 3ppm, and Ca = 270ppm were treated by adding calcium hydroxide thereto so 10 as to obtain prescribed pH values, and the analysis performed according to each pH value exhibited the results as shown in Table 6.

(Table 6)

pH	8	8.5	9.5	10	10.5	11
Ca (ppm)	100	36	26	20	25	65

15 As shown in Table 6, it can be seen that a removal rate of Ca is significantly improved in the range of pH = 8.5 - 105.

(Embodiment 6)

Four liters of waste water treated through the calcium carbonate packed tower exhibiting pH = 6.8, F = 3ppm, and Ca = 270ppm were treated by adding thereto 5.4g calcium hydroxide 20 of 96% purity so that pH value of the water becomes 10, and Table 7 shows the results of the analysis by sampling the water at each elapsed time.

(Table 7)

Reaction time (min)	5	15	30	60	120
Ca (ppm)	33	28	27	23	23

As shown in Table 7, the reaction is completed within a short time, and the water is sufficiently treated in 15 - 30 5 minutes. It can be seen from the table that the treatment is preferably carried out in 15 - 30 minutes because the effect is saturated even if it exceeds 30 minutes.

#### Industrial Utility:

According to the present invention, in the case of 10 effectively recover calcium fluoride by treating fluorine-containing waste water drained out of, for example, a semiconductor manufacturing plant with calcium carbonate, it is possible to reduce Ca concentration to a level to clear the water quality standards for industrial use by such a simple 15 chemical treatment as calcium carbonate is added to waste water dissolving calcium bicarbonate with 200-500ppm calcium concentration and is just stirred and reacted at ordinary temperatures, and this contributes to reuse of water resource.